

Chemistry
Higher level
Paper 2

Thursday 14 May 2015 (afternoon)

Candidate session number

2 hours 15 minutes

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Instructions to candidates

- Write your session number in the boxes above.
- Do not open this examination paper until instructed to do so.
- Section A: answer all questions.
- Section B: answer two questions.
- Write your answers in the boxes provided.
- A calculator is required for this paper.
- A clean copy of the **chemistry data booklet** is required for this paper.
- The maximum mark for this examination paper is **[90 marks]**.



Section A

Answer **all** questions. Write your answers in the boxes provided.

1. Ethanedioic acid is a diprotic acid. A student determined the value of x in the formula of hydrated ethanedioic acid, $\text{HOOC-COOH} \cdot x\text{H}_2\text{O}$, by titrating a known mass of the acid with a $0.100 \text{ mol dm}^{-3}$ solution of $\text{NaOH}(\text{aq})$.

0.795 g of ethanedioic acid was dissolved in distilled water and made up to a total volume of 250 cm^3 in a volumetric flask.

25 cm^3 of this ethanedioic acid solution was pipetted into a flask and titrated against aqueous sodium hydroxide using phenolphthalein as an indicator.

The titration was then repeated twice to obtain the results below.

Volume of $0.100 \text{ mol dm}^{-3} \text{ NaOH} / \text{cm}^3$	Titration 1	Titration 2	Titration 3
Final burette reading (± 0.05)	13.00	25.70	38.20
Initial burette reading (± 0.05)	0.00	13.00	25.70
Volume added			

- (a) Calculate the average volume of NaOH added, in cm^3 , in titrations 2 and 3, and then calculate the amount, in mol, of NaOH added.

[2]

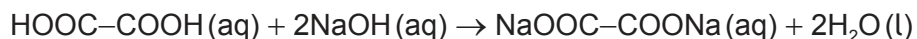
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(Question 1 continued)

- (b) (i) The equation for the reaction taking place in the titration is:



Determine the amount, in mol, of ethanedioic acid that reacts with the average volume of NaOH (aq). [1]

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- (ii) Determine the amount, in mol, of ethanedioic acid present in 250 cm³ of the original solution. [1]

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- (iii) Determine the molar mass of hydrated ethanedioic acid. [1]

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- (iv) Determine the value of x in the formula HOOC-COOH•xH₂O. [2]

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- (c) Identify the strongest intermolecular force in solid ethanedioic acid. [1]

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(Question 1 continued)

- (d) Deduce the Lewis (electron dot) structure of ethanedioic acid, $\text{HOOC}-\text{COOH}$. [1]

- (e) Predict and explain the difference in carbon-oxygen bond lengths in ethanedioic acid and its conjugate base, ${}^-\text{OOC}-\text{COO}^-$. [3]

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2. This question is about the compounds of some period 3 elements.

(a) State the equations for the reactions of sodium oxide with water and phosphorus(V) oxide with water. [2]

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(b) (i) Explain why the melting point of phosphorus(V) oxide is lower than that of sodium oxide in terms of their bonding and structure. [2]

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(ii) Predict whether phosphorus(V) oxide and sodium oxide conduct electricity in their solid and molten states. Complete the boxes with "yes" or "no". [2]

	Phosphorus(V) oxide	Sodium oxide
Solid state
Molten state

(This question continues on the following page)



(Question 2 continued)

- (c) Predict and explain the pH of the following aqueous solutions, using equations to support your answer.

[4]

Ammonium chloride, NH_4Cl (aq):

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Sodium methanoate, HCOONa (aq):

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3. The rate of reaction is an important factor in industrial processes such as the Contact process to make sulfur trioxide, SO₃(g).

(a) Define the term *rate of reaction*. [1]

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(b) Describe the collision theory. [3]

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(c) The Contact process involves this homogeneous equilibrium:



(i) State and explain how increasing the pressure of the reaction mixture affects the yield of SO₃. [2]

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(Question 3 continued)

- (ii) 2.00 mol of SO₂(g) are mixed with 3.00 mol of O₂(g) in a 1.00 dm³ container until equilibrium is reached. At equilibrium there are 0.80 mol of SO₃(g).

Determine the equilibrium constant (K_c) assuming all gases are at the same temperature and pressure.

[4]

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- (iii) State the effect of increasing temperature on the value of K_c for this reaction.

[1]

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- (d) Outline the economic importance of using a catalyst in the Contact process.

[2]

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4. Copper is a metal that has been used by humans for thousands of years.

(a) State the full electron configuration of ^{65}Cu . [1]

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(b) State one difference in the physical properties of the isotopes ^{63}Cu and ^{65}Cu and explain why their chemical properties are the same. [2]

Physical:
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Chemical:
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(c) Describe the bonding in solid copper. [2]

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Section B

Answer **two** questions. Write your answers in the boxes provided.

5. Ethanol has many industrial uses.

(a) State an equation for the formation of ethanol from ethene and the necessary reaction conditions.

[3]

Equation:

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Conditions:

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(b) (i) Define the term *average bond enthalpy*.

[2]

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(ii) Ethanol can be used as a fuel. Determine the enthalpy of combustion of ethanol at 298 K, in kJ mol^{-1} , using the values in table 10 of the data booklet, assuming all reactants and products are gaseous.

[4]

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(Question 5 continued)

- (c) Students can also measure the enthalpy of combustion of ethanol in the laboratory using calorimetry. Suggest the major source of systematic error in these procedures. [1]

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- (d) State the equation for the acid-catalysed reaction of ethanol with propanoic acid and state the name of the organic product. [2]

Equation:
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Name of the organic product:
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- (e) (i) A polyester can be formed when ethane-1,2-diol reacts with benzene-1,4-dicarboxylic acid.
Deduce the structure of the repeating unit and state the other product formed. [2]

Repeating unit:

Other product:
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(This question continues on the following page)



(Question 5 continued)

- (ii) State the type of polymerization that occurs. [1]

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- (f) (i) The standard enthalpy change of combustion, ΔH_c^\ominus , of propanoic acid is $-1527 \text{ kJ mol}^{-1}$. Determine the standard enthalpy change of formation of propanoic acid, in kJ mol^{-1} , using this information and data from table 12 of the data booklet. [4]

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- (ii) Deduce, giving a reason, the sign of the standard entropy change of the system for the formation of propanoic acid from its elements. [2]

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- (g) Identify **three** allotropes of carbon and describe their structures. [4]

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6. Bromomethane was used as a pesticide until it was found to be ozone-depleting.

(a) State the equation for the reaction between methane and bromine to form bromomethane. [1]

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(b) (i) Explain, using equations, the complete free-radical mechanism for the reaction of methane with bromine, including necessary reaction conditions. [4]

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(ii) Bromomethane reacts with aqueous sodium hydroxide. State the organic product of this reaction. [1]

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(c) Explain why the rate of the reaction between iodomethane, CH_3I , and $\text{NaOH}(\text{aq})$ is faster than the rate of the reaction between CH_3Br and $\text{NaOH}(\text{aq})$. [2]

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(Question 6 continued)

- (d) (i) Bromine can be produced by the electrolysis of **molten** sodium bromide. Deduce the half-equation for the reaction at each electrode. [2]

Positive electrode (anode):
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Negative electrode (cathode):
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- (ii) Predict the products formed at the electrodes during the electrolysis of concentrated **aqueous** sodium bromide. [2]

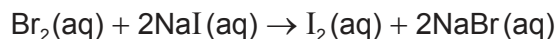
Positive electrode (anode):
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Negative electrode (cathode):
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(Question 6 continued)

- (e) Bromine reacts with aqueous sodium iodide.



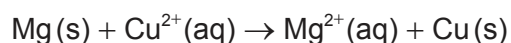
Identify the oxidizing agent in this reaction. [1]

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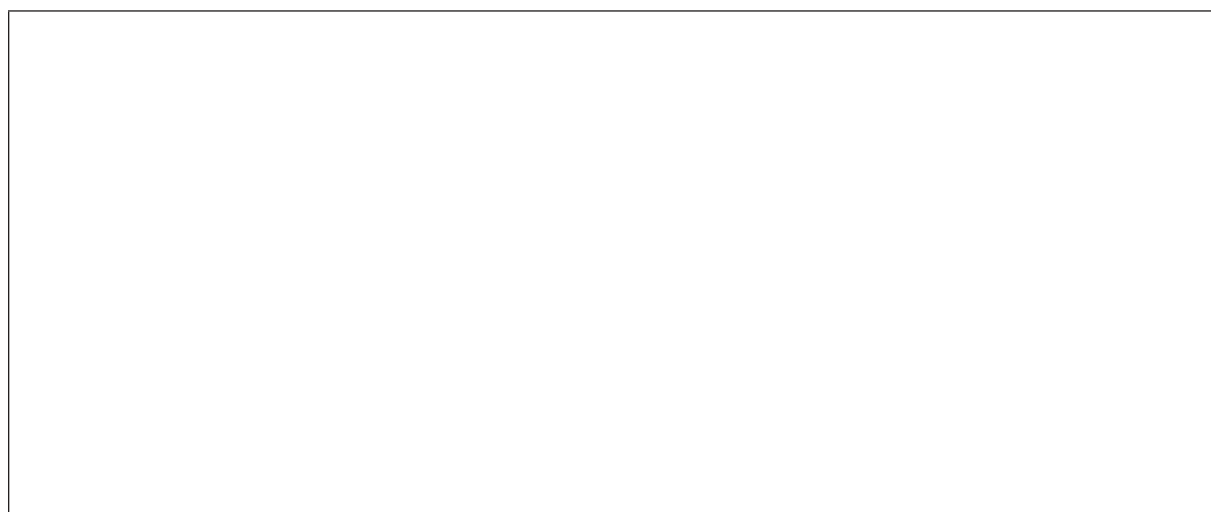
- (f) (i) Define the term *standard electrode potential*, E^\ominus . [1]

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- (ii) Draw a labelled diagram for the voltaic cell in which the following reaction occurs.



Include in your answer the direction of electron flow and the polarity of the electrodes. [4]



(This question continues on the following page)



(Question 6 continued)

- (iii) A student measures a voltage of 2.65 V in the voltaic cell formed between magnesium and copper half-cells using a digital voltmeter.

State the random uncertainty of this value, in V, and the number of significant figures in the answer.

[2]

Random uncertainty: Significant figures:

- (iv) Outline how the student can reduce the random error in her results.

[1]

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- (g) Determine the standard enthalpy change of formation, ΔH_f^\ominus , of NaCl(s), in kJ mol^{-1} , using a Born-Haber cycle and tables 7, 10 and 13 of the data booklet. The standard enthalpy change of atomization (standard enthalpy change of sublimation), $\Delta H_{\text{at}}^\ominus$, of Na(s) is $+108 \text{ kJ mol}^{-1}$.

[4]

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7. (a) (i) Ethanol is a primary alcohol that can be oxidized by acidified potassium dichromate(VI). Distinguish between the reaction conditions needed to produce ethanal and ethanoic acid. [2]

Ethanal:
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Ethanoic acid:
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- (ii) Determine the oxidation number of carbon in ethanol and ethanal. [2]

Ethanol:
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Ethanal:
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- (iii) Deduce the half-equation for the oxidation of ethanol to ethanal. [1]

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- (iv) Deduce the overall redox equation for the reaction of ethanol to ethanal with acidified potassium dichromate(VI). [2]

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(Question 7 continued)

- (b) Ethanol can be made by reacting aqueous sodium hydroxide with bromoethane. Explain the mechanism for this reaction, using curly arrows to represent the movement of electron pairs. [4]

- (c) (i) Determine the orders of reaction of the reactants and the overall rate expression for the reaction between 2-bromobutane and aqueous sodium hydroxide using the data in the table.

Experiment	[NaOH] / mol dm ⁻³	[C ₄ H ₉ Br] / mol dm ⁻³	Rate / mol dm ⁻³ s ⁻¹
1	1.00	1.00	1.66 × 10 ⁻³
2	0.50	1.00	8.31 × 10 ⁻⁴
3	0.25	0.25	1.02 × 10 ⁻⁴
4	1.00	0.50	8.29 × 10 ⁻⁴

[2]

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(Question 7 continued)

(ii) Determine the rate constant, k , with its units, using the data from experiment 3. [2]

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(iii) Identify the molecularity of the rate-determining step in this reaction. [1]

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(d) 2-bromobutane exists as optical isomers.

(i) State the essential feature of optical isomers. [1]

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(ii) Outline how a polarimeter can distinguish between these isomers. [2]

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(e) Describe, using an equation, the elimination of HBr from 2-bromobutane, stating the reagent used. [2]

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24EP19

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(Question 7 continued)

- (f) Describe the formation of σ and π bonds in an alkene. [2]

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- (g) The two most abundant isotopes of bromine have the mass numbers 79 and 81. Calculate the relative abundance of ^{79}Br using table 5 of the data booklet, assuming the abundance of the other isotopes is negligible. [2]

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8. Acids can be described as strong or weak.

- (a) (i) Outline the difference in dissociation between strong and weak acids of the same concentration. [1]

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- (ii) Describe **three** tests that can be carried out in the laboratory, and the expected results, to distinguish between $0.10 \text{ mol dm}^{-3} \text{ HCl (aq)}$ and $0.10 \text{ mol dm}^{-3} \text{ CH}_3\text{COOH (aq)}$. [3]

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- (b) Calculate the pH, using table 15 of the data booklet, of a solution of ethanoic acid made by dissolving 1.40g of the acid in distilled water to make a 500 cm^3 solution. [4]

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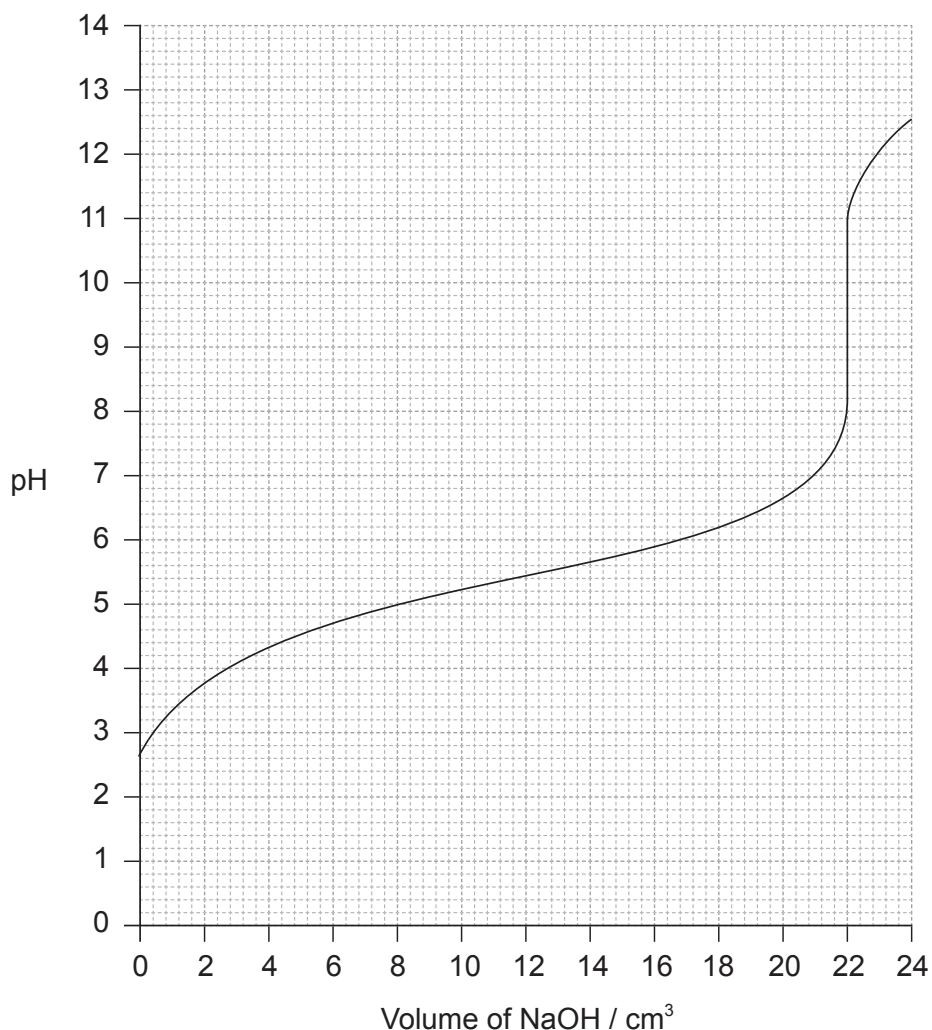
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(Question 8 continued)

- (c) (i) Determine the pH at the equivalence point of the titration and the pK_a of an unknown acid using the acid-base titration curve below.

[3]



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- (ii) Identify, using table 16 of the data booklet, a suitable indicator to show the end-point of this titration.

[1]

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(This question continues on the following page)



24EP22

(Question 8 continued)

- (iii) Describe how an indicator, that is a weak acid, works. Use Le Chatelier's principle in your answer. [2]

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- (d) (i) State the formula of the conjugate base of chloroethanoic acid, CH_2ClCOOH . [1]

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- (ii) Identify, with a reason, whether chloroethanoic acid is weaker or stronger than ethanoic acid using table 15 of the data booklet. [1]

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(Question 8 continued)

- (iii) Determine the pH of the solution resulting when 100 cm³ of 0.50 mol dm⁻³ CH₂ClCOOH is mixed with 200 cm³ of 0.10 mol dm⁻³ NaOH. [4]

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- (e) Describe how chlorine's position in the periodic table is related to its electron arrangement. [2]

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- (f) SCl₂ and SClF₅ are two sulfur chloride type compounds with sulfur having different oxidation states. Predict the name of the shape, the bond angle and polarity of these molecules. [3]

Molecule	Shape	Bond angle	Polarity
SCl ₂
SClF ₅

